Time: 3 Hours Max. Marks: 32

## SECTION - A

1C01CHE/PCH : Chemistry (For Physical and Biological Sciences)

Very short answer type. Each carry 1 mark. Answer all 5 questions.

- The force that allows electrons to move around the nucleus is \_\_\_\_\_\_
- 2. The hybridisation state of oxygen in water molecule is \_\_\_\_\_
- 3. To which block, the element with atomic number 56 belongs?
- 4. The solid outer part of the earth is known as \_\_\_\_\_\_
- AIBr<sub>3</sub> is formed by the reaction of a strong acid and a weak base. What is the nature of the solution, when this salt dissociates? (5×1=5)

## SECTION - B

Short answer type. Each carry 2 marks. Answer 4 questions out of 6.

- 6. What will be the uncertainty in the position of an electron (mass =  $9.10^{-31}$  kg) moving at 300 ms<sup>-1</sup> with an accuracy of 0.001 percent?
- 7. What is the Broglie wavelength of a tennis balls of mass 60 g moving with a velocity of 10 m/s is approximately : (Planck's constant,  $h = 6.63 \times 10^{-34} \, J \, s$ ).
- 8. What is bond order? Give bond order for B2 molecule.

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- 9. Explain BOD.
- 10. What are the harmful effects of SOx and NOx?
- 11. Write the Henderson equation for acidic buffer.

 $(4 \times 2 = 8)$ 

## SECTION - C

Short essay type. Each carry 3 marks. Answer 3 questions out of 5.

- 12. What is wave function? Give significance of  $\psi^2$ . Write Schrodinger equation.
- 13. What is lattice energy? Give Born Lande Equation explain the terms.
- 14. Give shape of BF3,CH4 and NH4+.
- 15. What are the causes of ozone depletion?
- 16. The molar solubility of PbBr $_2$  is 2.17  $\times$ 10<sup>-3</sup> M at a certain temperature. Calculate  $K_{\rm sp}$  for PbBr $_2$ . (3×3=9)

## SECTION - D

Long essay type. Each carry 5 marks. Answer 2 questions out of 4.

- 17. What are quantum numbers ? Explain all four quantum numbers.
- 18. What is hybridization? Explain sp3d and sp3d2 hybridization with examples.
- 19. Write notes on :
  - a) Global warming
  - b) Greenhouse effect.
- 20. Write notes on :
  - a) Common ion effect and solubility product.
  - b) pH of buffer solution.

(2×5=10)